

Application of Phase rule to one component system 7/9/2020 Page 1

From the phase rule, $F = C - P + 2$, it is clear that when $C = 1$, $P = 1$

$$F = 1 - 1 + 2 = 2$$

Thus all one component systems can be completely described graphically by stating two variables, pressure and temperature on appropriate axes.

One Component System ; Water System

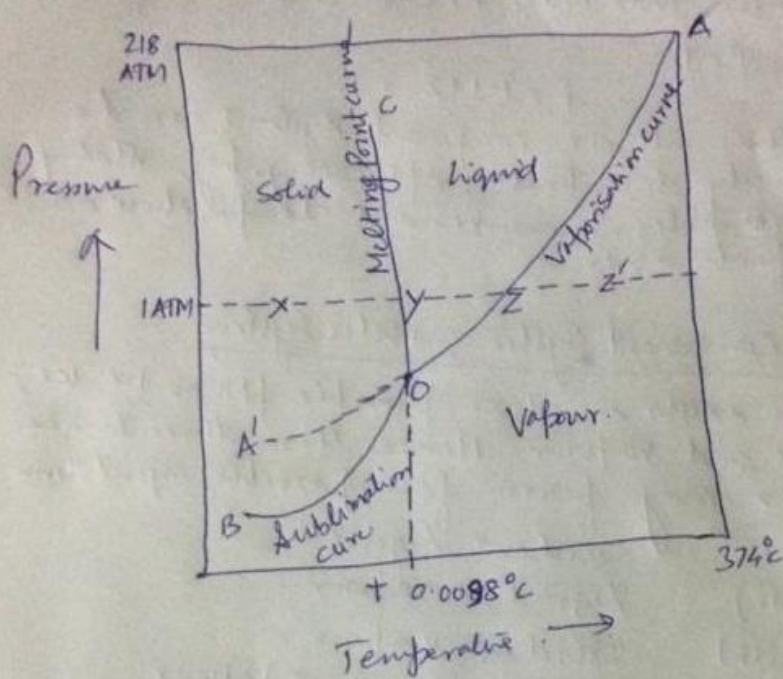
Water exists in three possible phases i.e. ice, water and vapour. Hence these three single phases may form four possible equilibria

- i) Solid \rightleftharpoons Liquid
- ii) Liquid \rightleftharpoons Vapour
- iii) Solid \rightleftharpoons Vapour
- iv) Solid \rightleftharpoons Liquid \rightleftharpoons Vapour

The phase diagram of water system is given in figure. The diagram consists of

- i) Three stable curves OB, OA and OC
- ii) One metastable curves OA'
- iii) Three areas AOB, COB and AOC
- iv) Triple point O

Phase diagram of water



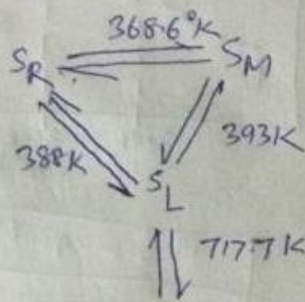
Table

System as given in the above diagram	Phase in equilibrium	Degree of freedom
<u>Lines</u>		
OA	Water \rightleftharpoons Vapour	1
OB	Ice \rightleftharpoons Vapour	1
OC	Ice \rightleftharpoons Water	1
OA'	Water \rightleftharpoons Vapour	1
<u>Areas</u>		
AOB	Vapour	2
COB	Ice	2
AOC	Ward water	2
<u>Triple point</u>		
O	Ice \rightleftharpoons Water \rightleftharpoons Vapour	0

One Component System : Sulphur System Page 3.

The Sulphur system is a one component system displaying polymorphism and solid-solid transformation. Sulphur exists in four phases

- i) Rhombic Sulphur (SR)
- ii) Monoclinic Sulphur (SM)
- iii) Sulphur Liquid (SL)
- iv) Sulphur vapour (SV)



According to phase rule $F = C - P + 2$, when $F = 0, C = 1$ then $0 = 1 - P + 2$ or $P = 3$. Thus all the four phases can not coexist in equilibrium.

The maximum number of phases which can coexist is three. The phase diagram of Sulphur system is given below

The phase diagram consists of

- i) Six stable curves, BO, OA, AE, OC, CF and AC.
- ii) Four metastable curves GO, OD, AD and DC.
- iii) Four Areas BOCF, BOAE, FC AE and OCA.
- iv) Triple point O, A, C and D.

Phase diagram Sulphur System

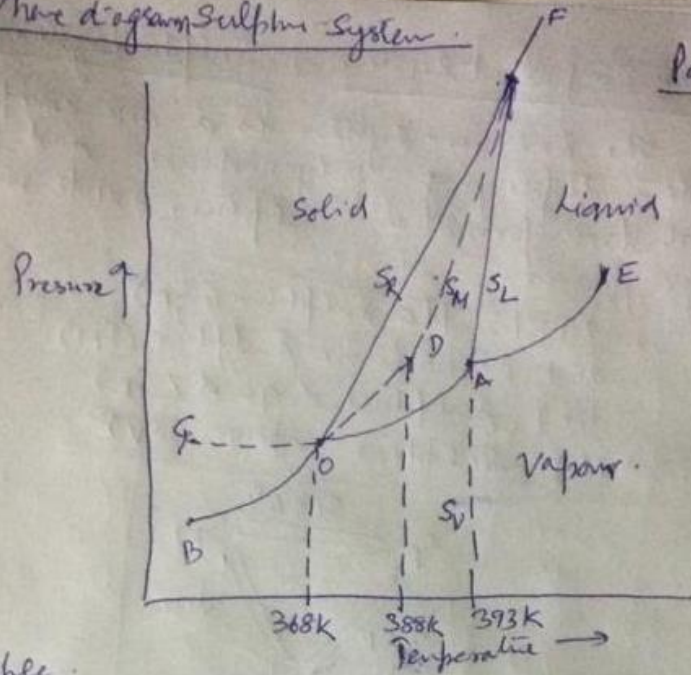


Table:

Name of System	Phase in equilibrium	Degrees of freedom
<u>Curves</u>		$F = C - P + 2$ $= 1 - 2 + 2 = 1$
i) BO(OD)	$SR \rightleftharpoons SV$	1
ii) OA(OG)	$SM \rightleftharpoons SV$	1
iii) AE(AD)	$SL \rightleftharpoons SV$	1
iv) DC(CF)	$SR \rightleftharpoons SL$	1
v) DC	$SR \rightleftharpoons SM$	1
vi) AC	$SM \rightleftharpoons SL$	1
<u>Areas</u>		$F = 1 - 1 + 2 = 2$
i) Below BOAE	SV	2
ii) To the right of EACF	SL	2
iii) Enclosed by OCA	SM	2
iv) To the left of BOCF	SR	2
<u>Triple Point</u>		$F = 1 - 3 + 2 = 0$
i) O	$SR \rightleftharpoons SM \rightleftharpoons SV$	0
ii) A	$SM \rightleftharpoons SL \rightleftharpoons SV$	0
iii) D	$SR \rightleftharpoons SM \rightleftharpoons SL$	0
iv) D	$SR \rightleftharpoons SL \rightleftharpoons SV$	0